

[REPLACEMENT SHEET]

A Compound Containing a Labile Disulfide Bond

This application is a continuation-in-part of serial no. 09/312,351 filed on May 14, 1999, pending, which claims the benefit of U.S. Provisional application No. 60/085,764, filed on May 16, 1998.

Background

Bifunctional molecules, commonly referred to as crosslinkers, are used to connect two molecules together. Bifunctional molecules can contain homo or heterobifunctionality. The disulfide linkage (RSSR') may be used within bifunctional molecules. The reversibility of disulfide bond formation makes them useful tools for the transient attachment of two molecules. Disulfides have been used to attach a bioactive compound and another compound (Thorpe, P.E. *J. Natl. Cancer Inst.* 1987, 79, 1101). The disulfide bond is reduced thereby releasing the bioactive compound. Disulfide bonds may also be used in the formation of polymers (Kishore, K., Ganesh, K. in *Advances in Polymer Science*, Vol. 21, Saegusa, T. Ed., 1993).

There are many commercially available reagents for the linkage of two molecules by a disulfide bond. Additionally there are bifunctional reagents that have a disulfide bond present. Typically, these reagents are based on 3-mercaptopropionic acid, i.e. dithiobispropionate. However, the rate at which these bonds are broken under physiological conditions is slow. For example, the half life of a disulfide derived from dithiobispropionimide, an analog of 3-mercaptopropionic acid, is 27 hours in vivo (Arpicco, S., Dosio, F., Brusa, P., Crosasso, P., Cautel, L. *Bioconjugate Chem.* 1997, 8, 327.). A stable disulfide bond is often desirable, for example when purification of linked molecules or long circulation *in vivo* is needed. For this reason, attempts have been made to make the disulfide less susceptible to cleavage.

It has been demonstrated that both stability, measured as reduction potential, and rate, measured as rate constants, of disulfide reduction are both related to the acidity of the thiols which constitute the disulfide. Additional factors that may affect the rate of reduction are steric interactions, and intramolecular disulfide cleavage. Looking at the